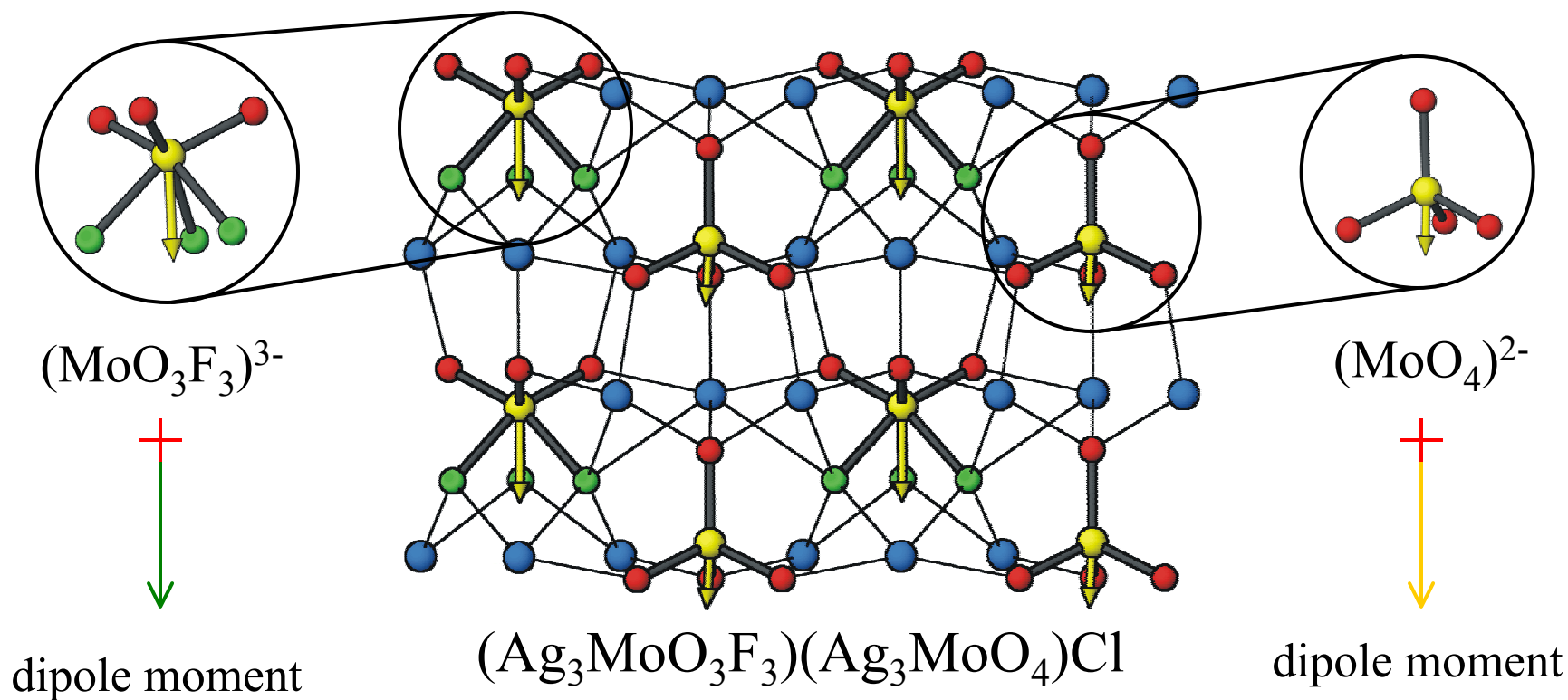


Solid State Oxyfluorides

Kenneth R. Poeppelmeier, Northwestern University, DMR-0312136



The combination of the polar $(\text{MoO}_3\text{F}_3)^{3-}$ and $(\text{MoO}_4)^{2-}$ “building blocks” creates a perfect alignment of the polar “anions” and leads to a noncentrosymmetric compound.

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The NSF Division of Materials Research award supports a comprehensive and interdisciplinary educational program in inorganic solid state chemistry.

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